

Question Number	Answer	Acceptable answers	Mark
<b>1(a)(i)</b>	fractional distillation		<b>(1)</b>

Question Number	Answer	Acceptable answers	Mark
<b>1(a)(ii)</b>	to make it liquid	liquefy/condense to remove water (vapour) to remove carbon dioxide	<b>(1)</b>

Question Number	Answer	Acceptable answers	Mark
<b>1(b)</b>	D weak forces of attraction between the oxygen molecules		<b>(1)</b>

Question Number	Answer	Acceptable answers	Mark
<b>1(c)(i)</b>	An description including <ul style="list-style-type: none"> <li>shared (electrons) (1)</li> <li>pair(s) of electrons (between atoms) (1)</li> </ul>	Ignore reference to complete/full shells Ignore reference to between two metals Ignore reference to between metal and non-metal Ignore reference to between molecules Any reference to between ions scores 0	<b>(2)</b>

Question Number	Answer	Acceptable answers	Mark
<b>1(c)(ii)</b>	2.4		<b>(1)</b>

Question Number	Answer	Acceptable answers	Mark
<b>1(c)(iii)</b>	diagram showing <ul style="list-style-type: none"> <li>any shared pair of electrons between a carbon and oxygen atom in CO<sub>2</sub> molecule (1)</li> <li>rest of molecule correct (1)</li> </ul>	Must have O C O arrangement If any atom labelled must be correct  Ignore inner electrons even if wrong electrons can be on/in ring or no ring Ignore intersecting circles  Accept all permutations of dots and crosses	<b>(2)</b>

Question number	Answer	Additional guidance	Mark
2(a)(i)	<ul style="list-style-type: none"> <li>particles are same size when they should be different sizes (1)</li> <li>model is in 2D but crystal is 3D (1)</li> </ul>	Allow reverse statements giving correct information.	(2)

Question number	Answer	Mark
2(a)(ii)	<p>An explanation that combines identification – knowledge (1 mark) and reasoning/justification – understanding (2 marks):</p> <ul style="list-style-type: none"> <li>very strong bonds/ionically bonded (1)</li> <li>between 2+ cations and 2- anions (1)</li> <li>so requires lot of energy to separate magnesium and oxide ions to melt the solid (1)</li> </ul>	(3)

Question number	Answer	Additional guidance	Mark
2(b)(i)	$\text{CaCO}_3 + 2\text{HCl} \rightarrow \text{CaCl}_2 + \text{H}_2\text{O} + \text{CO}_2$ <ul style="list-style-type: none"> <li>all formulae on correct side (2)</li> <li>balancing (1)</li> </ul>	Allow 3/4 formulae (1)	(3)

Question number	Answer	Additional guidance	Mark
2(b)(ii)	<p>relative formula mass copper carbonate  <math>= 63.5 + 12.0 + (3 \times 16.0)</math>  <math>= 123.5</math>            relative formula mass copper oxide  <math>= 63.5 + 16.0</math>  <math>= 79.5</math> (1)</p> <p>mass copper oxide  <math>= \frac{15.0 \times 79.5}{123.5} = 9.7 \text{ g to 2 s.f.}</math> (1)            Answer must be to two significant figures</p> <p>OR</p> <p>moles of copper carbonate  <math>= \frac{15.0}{123.5} = 0.12145</math> (1)            mass of copper oxide  <math>= \text{moles CuCO}_3 \times 79.5</math>  <math>= 9.7 \text{ g to 2sf}</math> (1)            Answer must be to two significant figures</p>	Award full marks for correct numerical answer without working.	(2)

Question number	Answer	Additional guidance	Mark
2(c)	<p>2.4/24 moles Mg = 0.1 mol (1)</p> <p>and 0.2 moles H<sub>2</sub>O has mass 0.2 × formula mass H<sub>2</sub>O = 3.6 g (1)</p> <p>total mass reactants = 2.4 + 3.6 = 6.0 g is the same as total mass products = 5.8 + 0.2 = 6.0 g (1)</p>	Award full marks for correct numerical answer without working.	(3)

Question Number	Answer	Acceptable answers	Mark
<b>3(a)(i)</b>	electrons		<b>(1)</b>

Question Number	Answer	Acceptable answers	Mark
<b>3(a)(ii)</b>	transition (metals/ elements)	transitional ignore transient	<b>(1)</b>

Question Number	Answer	Acceptable answers	Mark
<b>3(b)</b>	An explanation linking the following points <ul style="list-style-type: none"> <li>hydrogen chloride {soluble/dissolves} (in water) (1)</li> <li>forms hydrochloric acid (1)</li> </ul>	hydrogen chloride reacts with water	<b>(2)</b>

Question Number	Answer	Acceptable answers	Mark
<b>3(c)</b>	An explanation including <b>two</b> of the following points <ul style="list-style-type: none"> <li>(orange) colour due to bromine (1)</li> <li>chlorine displaces bromine (1)</li> <li>(because) chlorine is more reactive (than bromine) (1)</li> </ul>	chlorine displaces bromide (ions) a displacement reaction (occurs)OWTE	<b>(2)</b>

Question Number	Answer	Acceptable answers	Mark
<b>3(d)</b>	A description including <b>three</b> of the following points <ul style="list-style-type: none"> <li>mix solutions (1)</li> <li>filter (1)</li> <li>wash (precipitate / solid) with water (1)</li> <li>dry (precipitate / solid) in oven /leave to dry(1)</li> </ul>	pour (both) solutions into {beaker/other suitable container} ignore addition of hydrochloric acid  if wrong things mixed allow max 2 from last three points	<b>(3)</b>

Question Number	Answers	Acceptable Answers	Mark	
<b>4 (a)(i)</b>		chlorine-35	chlorine-37	
	number of protons	<b>17</b>	<b>17</b>	
	number of neutrons	<b>18</b>	<b>20</b>	
	number of electrons	<b>17</b>	<b>17</b>	
	the four 17s (1)			
	the 18 and 20 (1)			<b>(2)</b>

Question Number	Answers	Acceptable Answers	Mark
<b>4 (a)(ii)</b>	An explanation linking		
	<p>M1 average (mass of atoms/isotopes present) (1)</p> <p>M2 more chlorine-35 than chlorine-37 / higher {percentage / abundance} of Cl-35 / lower {percentage / abundance} of Cl-37 / (1)</p>	<p>mean ignore weight</p> <p>75% chlorine-35 / 25% chlorine-37/ chlorine-35 and chlorine-37 in ratio 3:1 / correct calculation to obtain 35.5 (2) eg <math>[(75 \times 35) + (25 \times 37)] / 100</math></p>	<b>(2)</b>

Question Number	Answers	Acceptable Answers	Mark
<b>4 (b)</b>	Diagram showing one carbon and four chlorines	use of dots or crosses or mixture of both	
	four pairs of electrons shared between the carbon and chlorine atoms (1)	ignore inner shells even if incorrect ignore symbols	
	fully correct (1)		<b>(2)</b>

Question Number	Indicative Content	Mark
<b>QWC</b>	<p data-bbox="331 242 400 275"><b>4(c)</b></p> <p data-bbox="435 242 1129 275">A response including some of the following points</p> <p data-bbox="435 307 1257 340">Note: (carbon to carbon) strong bonds is given in question</p> <p data-bbox="435 373 571 406">Diamond:</p> <p data-bbox="435 438 715 471">Uses and Properties</p> <ul data-bbox="483 504 895 711" style="list-style-type: none"> <li>• in cutting tools/engraving</li> <li>• drill bit</li> <li>• jewellery</li> <li>• diamond very hard/strong</li> <li>• attractive/lustrous</li> <li>• high melting point</li> </ul> <p data-bbox="435 777 616 810">Explanations</p> <ul data-bbox="483 843 1278 1017" style="list-style-type: none"> <li>• giant molecular/covalent</li> <li>• each carbon atom bonded to four other carbon atoms</li> <li>• three dimensional structure</li> <li>• to break it lots of bonds would need to be broken</li> <li>• would need lot of energy/force</li> </ul> <p data-bbox="435 1083 568 1116">Graphite:</p> <p data-bbox="435 1148 727 1181">Uses and Properties</p> <ul data-bbox="483 1214 922 1421" style="list-style-type: none"> <li>• to make electrodes</li> <li>• a lubricant</li> <li>• sporting equipment</li> <li>• in pencils/drawing</li> <li>• graphite conducts electricity</li> <li>• soft</li> </ul> <p data-bbox="435 1454 616 1487">Explanations</p> <ul data-bbox="483 1520 1294 1793" style="list-style-type: none"> <li>• giant molecular/covalent</li> <li>• each carbon atom bonded to three other carbon atoms</li> <li>• each carbon atom has a free electron</li> <li>• delocalised electrons</li> <li>• (delocalised) electrons move to carry current</li> <li>• layers of carbon atoms</li> <li>• weak forces/bonds between layers/sheets</li> <li>• so layers/sheets can slide/rub off or over each other</li> </ul>	<b>(6)</b>

Question number	Answer	Additional guidance	Mark
5(a)	An explanation that combines identification via a judgement (1 mark) to reach a conclusion via justification/reasoning (1 mark): <ul style="list-style-type: none"> <li>a negative ion must have more electrons than protons in the particle (1)</li> <li>therefore Z will have a 2- charge (1)</li> </ul>	Do not allow any comparison involving neutrons.	(2)

Question number	Answer	Additional guidance	Mark
5(b)	$40 + 2 \times (14 + 16 \times 3)$ (1) = 164 (1)	Award full marks for correct numerical answer without working.	(2)

Question number	Answer	Mark
5(c)	<ul style="list-style-type: none"> <li>Li ion with empty outer shell (1)</li> <li>1+ charge on Li (1)</li> <li>8 electrons on outer shell of F (1)</li> <li>1- charge on F (1)</li> </ul>	(4)